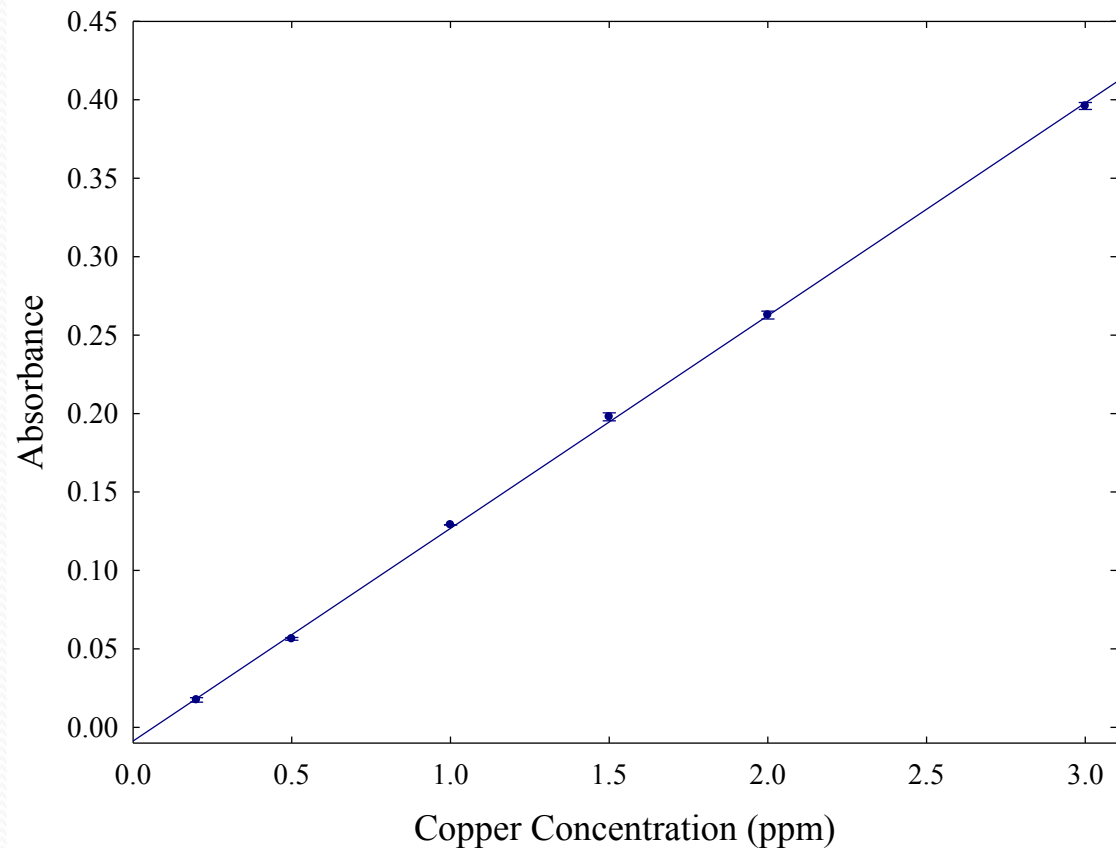


Calibration Curves

<u>[Cu] (ppm)</u>	<u>signal</u>
0.20	0.0175
0.50	0.0564
1.00	0.1290
1.50	0.1979
2.00	0.2628
3.00	0.3960

$$\text{sig} = -0.0087 + 0.1356[\text{Cu}]$$

If an unknown sample has a measured absorbance of 0.2549, determine the concentration of Cu in the sample.



$$[\text{Cu}] = \frac{(\text{sig} - [-0.0087])}{0.1356} = \frac{(0.2549 + 0.0087)}{0.1356} = 1.944 \text{ ppm}$$