Amedeo Avogadro (1776-1856) was an Italian chemist. Though not immediately accepted by other scientists, he clearly distinguished atoms and molecules from one another and stated that gases are composed of molecules which are composed of atoms. The number of particles in one mole of a substance was named after him. --Alanna

Avogadro was an Italian savant known for Avogadro's law and number. His law says that equal amounts of gas at STP contain the same amount of particles and his number is the amount of atoms in 12 grams of carbon-12. Avogadro distinguished between molecules and atoms and mass and weight. --Andrew

Amedeo Avogadro (Italy, 1776-1856) is known for Avogadro’s number (6.02 x 10\(^{23}\) items in a mole) and Avogadro’s law, which states that equal volumes of gases will have equal numbers of molecules. He acknowledged that substances were made of molecules which were in turn made of “elementary molecules”, or atoms. --Chantal

The Italian chemist Amadeo Avogrado lived from 1776-1856 in Europe. In explaining Gay-Lussac’s Law of Combining Gases and reinforcing Dalton’s atomic theory, Avogadro remarked that the atoms, in whole number relationships, could travel in pairs. He is best known for his number 6.02 X 10\(^{23}\). --Christine

Born in 1776 Italy, Avogadro is best known for his number, 6.02 x 10\(^{23}\), which expresses the amount of “elementary entities” in one mole of a substance. He also proposed Avogadro’s law, stating that gases with equal volume in the same pressure and temperature have the same number of molecules. --Davy

Amende Avogadro born in 1776 in Italy was the first to distinguish between a molecule and an atom and to come up with a constant (6.02 x 1023) for which tells us how many things are in a mole. Avogadro also concluded that there were 3 “molecules”, the three phases. --Denise

Amadeo Avogadro, an Italian chemist, lived from 1776 to 1856 and created a number used to mass atoms. He discovered that equal volumes of gases at the same temperature and pressure have the same number of molecules. Avogadro distinguished atoms from molecules, allowing people to understand chemical reactions better. --Eric
Amadeo Avogadro was born on August 9, 1776 and died on July 9, 1856. He is most commonly known for Avogadro’s number, which is the amount of protons or neutrons in one gram. He is also known for finding out about molecules and diatomic atoms. --Eyad

Amadeo Avogadro, a physicist, lived in Italy from 1776 to 1856. By combining Gay-Lussac’s work and Dalton’s atomic theory, Avogadro published a paper regarding Avogadro’s principle that clearly showed the difference between a molecule and an atom. This helped the chemists to determine the molar masses and the atomic masses. --Isaac

Amadeo Avogadro (1776-1856) was an Italian physicist. He stated that masses of different gases with known volumes correspond to their molecular weights, so their molecular masses can be found from volume. He developed the unit mole, which contains 6.02x10²³ units; future chemists could measure the amount of units in a substance. --Jaya

Amadeo Avogadro (Italy, 1776 – 1856) is most famous for Avogadro’s number (6.022 x 10²³), the number of molecules in one mole of a substance. Avogadro is also famous for Avogadro’s law, a gas law that states that the ratio between volume and moles of gas is constant. --Jeff

(1776-1856) Avogadro did not actually come up with the number 6x10²³, it was derived from his principle. Avogadro's principle states that equal volumes of all gases at the same temperature and pressure contain the same number of molecules. So, all substances have the same molecular weight in one mole. --Julia

Amadeo Avogadro (1776-1856) was an Italian savant. He was the person who created the number of elementary entities in 1 mol of a substance, 6.02 x 10²³. His greatest contributions to molecular theory, was Avogadro’s law. Because of his contributions he was hailed as a founder of atomic-molecular theory. --Karen

Lorenzo Romano Amedeo Carlo Avogadro was an Italian chemist who lived from 1776 to 1856. He is best known for Avogadro’s number (6.02 x 1023 items= 1 mole). He furthered Gay-Lussac’s studies to go beyond volumes of elements to the actual particles. He also studied elements with diatomic tendencies. --Katy

Amadeo Avogadro was an 19th century Italian chemist most recognized by Avogadros’s constant and his work in molecular chemistry. His famous Avogradros’s Law hypothesizes that the volume of any gas contains the same number of particles if it has the same temperature and pressure regardless of molecule size. --Madeline

Amadeo Avogadro was born in the year 1776 in Turin Italy. His most notable achievement was distinguishing between molecules and atoms. Contrary to popular belief, Avogadro did not find Avogadro’s number, however, he conducted research that enabled it to be found. --Oliver

Italian savant, Amadeo Avogadro, was best known for his contributions to molecular theory and Avogadro’s law which measures the relationship between gasses and their molecular weight. The law suggests that the amount of atoms, or molecules in a substance, equals Avogadro’s constant, which is known as 6.02x10^23. --Tokunbo
Amedeo Avogadro was an Italian physicist who acknowledged the distinction between atoms and molecules. He formulated a law recognizing the relationship between the volumes of gases and the amount of molecules prevalent in the sample. He also defined one mole of a given substance to be $6.02214179 \times 10^{23}$ molecules. --Victoria

Amadeo Avogadro, an Italian, lived from August 9, 1776 – July 9, 1856. One of his achievements was in clearing up the difference between a molecule and an atom. In one of his articles, he distinguished between "simple" molecules and molecules. He also is known for Avogadro's law and Avogadro's number. --Yulan