

Chemistry 101 Final Exam Guide for Chem 101-02 (DO NOT LIMIT YOUR REVIEW TO THESE QUESTIONS!) **Bring a scantron for the final exam.**

The 200 pt test is on Wednesday, Dec 3 at 8-10:30 am. The final exam is cumulative and will assume you have reviewed chapters 1-8 with greater emphasis placed on material after chapter 6.

For chapters 1-6, please refer also to the review guides for those two exams. As before, we will skip the sections involving oxidation-reduction reactions, and leave that for chem 103. We will also skip molecular orbital theory and leave that for next quarter.

Chapt 7. Quantum Theory:

- Do calculations involving photon energy, frequency and wavelength for electromagnetic radiation.
- Do calculations involving the energy levels of the hydrogen atom. What is the balmer series and how is it different from the Lyman and Paschen series?
- Know the rules for allowable quantum numbers of the electron's wave function.
- Be able to write the electron configuration of elements and their ions.(also using core notation).
- Be able to explain the periodic trends of: atomic radii, ionic radii, electron affinity and ionization potential

Chapt 8. Covalent Bonding:

- Be able to draw the Lewis structure for a molecule. Know the rules for drawing Lewis structures. (there are many examples in the book which you should try. Then compare your structure with the one given in the book so you can sharpen your skills. Determine the polarity of certain bonds using electronegativity.
- Know the various types of isomers possible.
- use bond enthalpies to determine the enthalpy of reaction.
- know how to determine formal charge to differentiate between possible Lewis structures.
- Know: resonance, octet rule.

Below are some practice review question. Be able to do all homework in chapt 8.

1) An acidic compound composed of 2.1%H, 29.8% N and 68.1% O has a molecular mass of 47 g/mol. a) What is the empirical formula of the compound? b) What is the name of the compound? c) What is the Lewis structure if H is bonded to O? d) What is the electron domain geometry around N? e) What is the molecular geometry of the molecule (describe all angles)? f) Give the formal charge of nitrogen and oxygen in the molecule. g) Predict the polar/nonpolar nature of this molecule? h) What are the hybrid orbitals present in the N atom? in the O atom?

2) Explain the periodic trends: atomic radii, ionization energy, electron affinity. Explain "anomalies" in the ionization energy trend involving 2nd row elements.

3) a) Using "box" notation, draw the occupied orbitals of neutral vanadium. How many unpaired electrons is in V? b) Explain why nitrogen has a higher first ionization energy than oxygen. Explain why it has a higher first ionization energy than carbon.

4) Periodic trends: Arrange following in order of a) increasing size: Ar, S²⁻, K²⁺, b) increasing ionization energy: F, S, Al, He c) increasing electronegativity: Se, Ne, O d) molecular polarity: H₂O, CO₂, NO₂⁻. Draw their Lewis structures.

5) Consider the following molecules, and fill in the information requested below each molecular formula:(note that some may violate the octet rule) H_2O , XeF_4 , PCl_5 , SO_4^{2-}

6) SKIP THE QUESTIONS WHICH ARE RELATED TO MO THEORY. Write down the electron configurations of Li_2 , O_2 , N_2 in terms of its MOs (molecular orbitals). Compare their bond orders and predict their magnetic properties. Write down their Lewis structures.

7) Consider the reaction of oxalic acid ($\text{H}_2\text{C}_2\text{O}_4$) with nitrous acid (HNO_2) to form carbon dioxide gas, nitrogen monoxide, and water. a) write the balanced equation. b) what type of reactions is this? If redox, identify the reducing and oxidizing agent. c) Determine the oxidation numbers of all elements involved.

8) Describe or explain scientific contributions made by Dalton, Mendeleev, Planck, Einstein, Thomson, Bohr, Schrodinger, de Broglie & Heisenberg in our modern understanding of the atom. What is the wavelength of a proton (1.67×10^{-27} kg) travelling at 1.2×10^5 m/s?

9) The energy required to convert O_2 molecules to O atoms is 496 kJ/mol. If electronic radiation of 180 nm is absorbed by 1 mole of O_2 molecules, how much kinetic energy will be present in the O atoms? What wavelength photons are required to "split" O_2 molecules to O atoms?

10) a) Name the following when pure and when in an aqueous solution: i) HClO_3 , ii) HClO_2 , & iii) H_2S :Draw their Lewis structures and identify polar bonds if any. b) Name the compounds: K_3PO_4 ; BaCl_2 ; Give the formulas for aluminum dichromate, magnesium phosphate

11) Give the number of electrons, protons and neutrons in: ferrous ion; mercuric ion or $^{195}_{78}\text{Pt}^{4+}$.. Write down the corresponding electron configurations for these ions.

12). The density of a hydrochloric acid (HCl) solution is 1.19 g/mL. If its concentration is 37% (mass percent). How many mLs of the solution would contain 3.01×10^{22} molecules of hydrogen chloride?

13) Consider the reaction of iron(II) chloride with potassium permanganate in an acidic solution. The balanced net ionic equation is given below: $5 \text{Fe}^{2+}(\text{aq}) + \text{MnO}_4^- (\text{aq}) + 8\text{H}^+(\text{aq}) \rightarrow 5 \text{Fe}^{3+}(\text{aq}) + \text{Mn}^{2+}(\text{aq}) + 4 \text{H}_2\text{O}$

a) Suppose you titrate 35.0 mLs of iron(II) chloride solution with 0.0500 M potassium permanganate, and the equivalence point is reached at 24.0 mLs, what is the original concentration of the iron(II) chloride solution?

b) What is the concentration of the iron(II) chloride in the solution 10.0 mLs before the equivalence point?