

You may keep this question sheet. Follow correct seating arrangement.

Chemistry 151

Lecture Quiz #2

Name KEY

1/20/05

25 points - 10 minutes

CIN _____

Multiple choice (Choose BEST answer): 2.5 points each. Use SMALL scantrons.

b) (1) Consider the formation of an ionic compound: $2K(s) + F_2(g) \rightarrow 2KF$. Choose the true statement:

b) The Fluorine (F) atom has become more negative.

c) (2) Consider the following elements: aluminum, fluorine, boron and helium. The order of **increasing atomic radii** for these elements is: **c)** He, F, B, Al

a) (3) Consider the following elements: aluminum, fluorine, boron and helium. The order of **increasing ionization energy** (I.E.) for these elements is: **a)** Al, B, F, He

a) (4) Consider the following elements: aluminum, fluorine, boron and helium. The order of **increasing electron affinity** for these elements is: **a)** Al, B, F, He

not counted (5) Consider the following elements: aluminum, fluorine, boron and helium. The order of **increasing metallic character** for these elements is: (*typo. it should have read decreasing instead of increasing; none of the answers is correct*)

c) (6) Which of the following dot structures is correct for Boron?

e) (7) Which of the following names would apply to phosphorus?

a) alkali metal b) alkaline earth c) halogen d) noble

b) (8) Which of the following is the correct way to write the ion for magnesium? **e)** none of the previous.

b) (9) Which of the following correctly refers to the ion of iron containing the lowest charge:

b) iron (II)

d) (10) The correct formula for aluminum carbonate is:

d) $Al_2(CO_3)_3$

c) (11) The correct name for the PO_4^{3-} is: **c)** phosphate.

b) (12) Consider the molecule carbon tetrafluoride. Choose the correct match of structure and polarity: